

# Chapter 5 Molecules And Compounds

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5.1: Sugar and Salt Both salt and sugar have radically different properties (both physical and chemical) than the constituent elements that make up these compounds. That is a central feature of chemical reactions as this chapter will discuss. 5.2: Compounds Display Constant Composition

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Ternary compounds are composed of three or more elements.

5.8: Naming Molecular Compounds Molecular compounds are inorganic compounds that take the form of discrete molecules. Examples include such familiar substances as water and carbon dioxide. These compounds are very different from ionic compounds like sodium chloride.

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Chapter 5 Molecules and Compounds. STUDY. PLAY. The properties of a compound are, in general, \_\_\_\_\_ from the properties of the elements that compose it. different. law of constant composition. All samples of a given compound have the same proportions of their constituent elements.

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Chapter 5: Molecules and Compounds Read Chapter 5 Check for MasteringChemistry due dates. Pure Substances and Mixtures: Pure substances have one invariable composition (elements and compounds) Mixtures have a variable composition. Homogeneous mixtures (solutions) mix

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## **Chapter 5: Molecules and Compounds - Moorpark College**

Chapter 5: Molecules and Compounds. STUDY. PLAY. Acid. A molecular compound that dissolves in solution to form  $H^+$  ions. They have the ability to dissolve some metals and will turn litmus paper red. Atomic element. An element that exists in nature with single atoms as the basic unit.

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Chapter 5: (Molecules and Compounds) STUDY. PLAY. Law of Constant Composition. All samples of a given compound have the same proportions of their constituent elements. Example of Mass ratio ( $H_2O$ ) 16.0 g O/ 2.0g H= 8.0 or 8.0:1. Chemical Formula.

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Chapter 5 Molecules and Compounds. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. deborah70601. Terms in this set (26) In a compound, the elements combine in fixed, definite proportions. Free atoms are rare in nature In a mixture, they can have any proportions whatsoever.

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Chapter 4: Atoms and Elements. Section 1: Early Ideas about Matter; Section 2: Atomic Discoveries; Section 3: Subatomic Particles and Elements; Section 4: The Periodic Table; Section 5: Ions; Section 6: Isotopes; Chapter 5: Molecules and Compounds. Section 1: Compounds and Chemical Formulas; Section 2: Writing Formulas for Ionic Compounds

## **Marnik, Jennifer / Chapter 5: Molecules and Compounds**

Molecules are groups of atoms that behave as a single unit.

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Some elements exist as molecules: hydrogen, oxygen, sulfur, and so forth. There are rules that can express a unique name for any given molecule, and a unique formula for any given name.

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MarcoChem7216. Nomenclature for ionic compounds type I and II, covalent compounds and acids. Terms in this set (14)

Compound. a grouping of bonded atoms that forms a molecule or formula unit. The elements that make up a particular ...

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Chapter 5: Molecules and Compounds Chapter 5 Page 1  
Formulas use element symbols and subscripts to the right of the element symbol to describe the number of atoms of each element in the compound. symbol for carbon and implied subscript of "1"

## **Chapter 5: Molecules and Compounds**

Chapter 5 Molecules and Compounds . Compounds Display Constant Composition Joseph Proust (1754-1826) The law of constant composition states: All samples of a given compound have the same proportions of their constituent elements. What does the law of constant composition mean?

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## **Ch 5 Molecules and Compounds**

Tro's "Introductory Chemistry", Chapter 5 5. How to Represent Compounds. • A compound is substance that is composed of atoms of two or more elements. • Compounds can be represented. Writing the symbol of each element present in the compound. describing the number of each atom in the simplest unit of the compound.

## **Chapter 5 Molecules and Compounds**

Compound: a pure substance that contains... Chemical formula: ratio of elements Law of constant composition: ratio in a compound is consistent if the compound is pure Formula unit: atoms represented by a chemical formula One formula unit of  $\text{Ca}_3(\text{PO}_4)_2$  contains... Chapter 5: Molecules and compounds ch5a Page 1

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## **Chapter 5: Molecules and compounds**

Introductory Chemistry (5th Edition) answers to Chapter 5 - Molecules and Compounds - Exercises - Problems - Page 156 28 including work step by step written by community members like you. Textbook Authors: Tro, Nivaldo J. , ISBN-10: 032191029X, ISBN-13: 978-0-32191-029-5, Publisher: Pearson

## **Chapter 5 - Molecules and Compounds - GradeSaver**

Normal hemoglobin is a tetramer, consisting of two molecules of  $\beta$  hemoglobin and two molecules of  $\alpha$  hemoglobin. In sickle-cell disease, as a result of a single amino acid change, the mutant hemoglobin tetramers associate with each other and assemble into large fibers.

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