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2) For the silver, calculate the mass and number of moles that was produced in this reaction. 3) Find the following ratios: a) The ratio of moles of Copper reacted to moles of silver nitrate reacted. b) The ratio of moles of silver nitrate reacted to moles of silver produced. c) The ratio of moles of copper reacted to moles of silver produced.

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O to moles (not on computer do this one on your own) 16. Convert 77.2 grams of CH 4 to molecules. (Show all work) 17. Convert 22.7 grams of H 2 O to molecules (not on computer do this one on your own) 18. What does the name mole actually come from? ... The Mole Webquest Author:

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Solved Examples on the Mole Concept. Some solved example questions on the mole concept are provided in this subsection. Q.1: How many moles of iron are present in a pure sample weighing 558.45 grams? A.1: The molar mass of iron is 55.845 g/mol. Therefore, the number of moles of iron in the pure sample weighing 558.45 grams is:

### Mole Concept - What is a Mole? [Related Formulae, Examples]

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